



(19)

Europäisches Patentamt
European Patent Office
Office européen des brevets



(11)

EP 1 007 472 B1

(12)

EUROPEAN PATENT SPECIFICATION

(45) Date of publication and mention
of the grant of the patent:
03.09.2003 Bulletin 2003/36

(51) Int Cl.7: **C01B 3/32, H01M 8/06**

(21) Application number: **99919449.1**

(86) International application number:
PCT/GR99/00020

(22) Date of filing: **19.05.1999**

(87) International publication number:
WO 99/061369 (02.12.1999 Gazette 1999/48)

(54) PROCESS FOR THE PRODUCTION OF HYDROGEN AND ELECTRICAL ENERGY FROM REFORMING OF BIO-ETHANOL

VERFAHREN ZUR HERSTELLUNG VON WASSERSTOFF UND ELEKTRISCHER ENERGIE
DURCH REFORMIEREN VON BIOETHANOL

PROCEDE DE PRODUCTION D'HYDROGÈNE ET D'ENERGIE ELECTRIQUE A PARTIR DU
REFORMAGE DE BIO-ETHANOL

(84) Designated Contracting States:
**AT BE CH CY DE DK ES FI FR GB GR IE IT LI LU
MC NL PT SE**

(30) Priority: **22.05.1998 GR 98100180**

(43) Date of publication of application:
14.06.2000 Bulletin 2000/24

(73) Proprietors:
• Kadmos AE Ependytikon Protovouliou
15125 Athens (GR)
• Verykios, Xenophon E.
Maroussi, 15125 Athens (GR)

(72) Inventor: **VERYKIOS, Xenophon
GR-264 41 Patra (GR)**

(74) Representative: **Argyriadis, Korinna
Sina Street 14
106 72 Athens (GR)**

(56) References cited:

- BIRDSELL S A WILLMS R S ET AL: "PURE HYDROGEN PRODUCTION FROM OCTANE, ETHANOL, METHANOL, AND METHANE REFORMING USING A PALLADIUM MEMBRANE REACTOR" IECEC '97. PROCEEDINGS OF THE 32ND INTERSOCIETY ENERGY CONVERSION ENGINEERING CONFERENCE. ENERGY SYSTEMS, RENEWABLE ENERGY RESOURCES, ENVIRONMENTAL IMPACT AND POLICY IMPACTS ON ENERGY, HONOLULU, HI, JULY 27 - AUG. 1, 1997, vol. 3 & 4, 27 July 1997 (1997-07-27), pages 1942-1946, XP000779068 AND ELECTRONICS ENGINEERS
- S. P. KOTHARI ET AL.: "Economics of biomass-derived alcohol application in fuel cells" ENERGY PROGRESS, vol. 2, no. 4, December 1982 (1982-12), pages 187-191, XP002102013 USA
- DATABASE APIIT [Online] THE AMERICAN PETROLEUM INSTITUTE, NEW YORK, NY, US document no. 4502054, 1998 F. HAGA ET AL.: "Effect of particle size on steam reforming of ethanol over alumina-supported cobalt catalyst" XP002102015 & NIPPON KAGAKU KAISHI, no. 11, November 1997 (1997-11), japan
- DATABASE WPI Section Ch, Week 8813 Derwent Publications Ltd., London, GB; Class E36, AN 88-088358 XP002112855 & JP 63 039626 A (YAMAHA HATSUDOKI KK), 20 February 1988 (1988-02-20)

- R. KUMAR ET AL.: "Fuel processing requirements and techniques for fuel cell propulsion power" ACS 206TH NATIONAL MEETING, vol. 38, no. 4, 1993, pages 1471-1476, XP002102014 USA
- PATENT ABSTRACTS OF JAPAN vol. 009, no. 045 (C-268), 26 February 1985 (1985-02-26) & JP 59 189935 A (MITSUBISHI JUKOGYO KK), 27 October 1984 (1984-10-27)

Description

[0001] The present invention refers to a process for the production of hydrogen and electrical energy from reforming of bio-ethanol, with the use of fuel cells and zero emission of pollutants. Ethanol is produced from biomass which contains sugar and/or cellulosic components, originating from any source. Aqueous solution of ethanol (40-70% by weight) is mixed with air (0-0,5 mol oxygen per mol ethanol) and is fed to a reactor which contains suitable catalyst so as for the reactions of partial oxidation and reforming of ethanol to take place. In the same or a different reactor the shift reaction for the consumption of carbon monoxide and further production of hydrogen is taking place.

[0002] The gaseous mixture which is produced in this manner is rich in hydrogen which can be separated and used in different applications. Alternatively, the gaseous mixture is fed to a fuel cell, preferably of the phosphoric acid or proton exchange membrane or solid polymer, in which electrical energy and heat are produced.

[0003] No emissions harmful to the environment are produced in any of the stages of the above process.

Technical field

[0004] The use of biomass as a renewable energy source has been investigated and proposed for many years [1]. Three different means of use of biomass as energy source have been applied internationally: Combustion, pyrolysis for the production of gaseous and liquid fuels, and fermentation for the production of ethanol. Sources of biomass can be plants which have certain specific characteristics and which are grown for this purpose, or waste materials from cultivation of edible products or from agroindustries or from forestry. Studies which have been conducted in recent years show that there are significant quantities and sources of biomass which can be utilized for the production of energy.

[0005] The production of ethanol from biomass which is often referred to in the bibliography as "bio-ethanol", is known and is practiced in large scale, mostly in North and South America and in Europe [2]. The processes for the production of ethanol can be classified in two large categories: Those which utilize sugar-containing raw materials - products of energetic cultivations (for example sweet sorgum) and those which utilize cellulosic raw materials originating from energetic cultivations (sorgum, cane, solid residue of sweet sorgum, etc) as well as from residues of agroindustries. In the first case sugars are directly fermented for the production of ethanol while in the second case the hydrolysis step or other processes are preceeded for the production of sugars which are then converted to ethanol via fermentation [1,2].

[0006] In R. KUMAR ET AL.: "Fuel processing requirements and techniques for fuel cell propulsion power" ACS 206TH NATIONAL MEETING, vol. 38, no. 4, 1993, pages 1471-1476, is described the reforming of primarily methanol with respect to catalysts and types of reformers, for application in fuel cell powered electrical vehicles. Reference is also made to an experiment concerning steam reforming of ethanol, which was not very successful since the efficiency towards hydrogen production was very small.

[0007] Although the technology of biomass fermentation for the production of ethanol is mature, it has not been applied in large scale, at least in Europe, for economic reasons. A large fraction of the cost of bio-ethanol is the cost of separation of the aqueous solution which derives from fermentation and which contains approximately 8 to 12% ethanol. In order to use ethanol as a fuel in internal combustion engines the required purity exceeds 99%. Because ethanol and water form an azeotropic solution when the ethanol content is approximately 95%, further purification requires energy-consuming techniques, the application of which increases significantly the cost of bio-ethanol.

Brief description of the invention

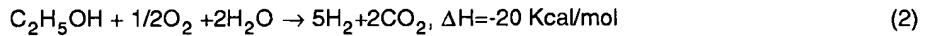
[0008] Purpose of the present invention is to utilize ethanol which is produced from biomass for the production of energy without the requirement of its separation from water to a large degree. With the new method which is presented, the cost of ethanol production is reduced significantly while, simultaneously, the thermodynamic efficiency of its use is increased significantly (with the application of fuel cells) and the gaseous pollutants which are produced during its burning (for example in internal combustion engines) are eliminated.

[0009] According to the present invention, a mixture of ethanol, oxygen and water react and produce hydrogen and carbon dioxide, and the hydrogen is fed to a fuel cell which produces electrical energy from the electrochemical oxidation of hydrogen to water. The reforming of ethanol with water which is described by the reaction:



has been little investigated, in contrast to the reforming of methanol. It has been reported that reforming of ethanol

over a copper catalyst of a Pd/ZnO produces acetic acid, acetaldehyde, H₂ and heavier oxygen-containing products [3]. Acetaldehyde is also produced when the reaction is taking place on Ni and Pt catalyst supported on MgO[4]. Complete reforming of ethanol on Ni catalyst has been reported, which, however, deactivates rapidly, probably due to carbon deposition [5]. The reaction of reforming of ethanol with water (Reaction 1) is endothermic and consequently supply of heat to the reactor is required. The simplest way to achieve this is the co-feeding of oxygen (air) together with the ethanol-water mixture over a suitable catalyst so as a small fraction of ethanol to be oxidized towards CO₂ and H₂O, producing the required heat for the reforming reactions:



[0010] As far as we know, no study has appeared in the literature referring to the partial oxidation/reforming of ethanol.

[0011] According to the invention which is presented the hydrogen which is produced from the partial oxidation/reforming of ethanol, is fed to a fuel cell for the production of electrical energy. The advantages of fuel cells are the zero emission of pollutants since the only product of combustion of hydrogen is water and the improved thermodynamic efficiency in comparison with internal combustion engines.

[0012] Fuel cells are an innovation which finds applications in the production of electrical energy without the emission of pollutants, consuming hydrogen as fuel which is oxidized electrochemically with oxygen, with simultaneous production of electrical energy [6]. The efficiency of fuel cells is approaching 70% of the heat which corresponds to the combustion of hydrogen, i.e. it is twice the efficiency of thermal engines which is subjected to the thermodynamic constraints of the Carnot type. There are five types of fuel cells which differ in the type of electrolyte and in the temperature of operation. 1) Phosphoric Acid which is the most commercially advanced type and operate at temperatures around 200°C. They produce electrical energy with efficiency from 40 to 80% [7]. 2) Proton exchange membranes or solid polymers which operate at low temperatures, approximately 100°C, and offer large power density with quick response to power demands [8]. 3) Solid Oxide Fuel Cells which can be utilized in applications in which large power is required and operate at about 1000°C [9]. 4) Molten Carbonate Salts, whose electrolyte consists of molten carbonate salts of Li and K, the fuel is hydrogen and CO and the oxidizing mixture consists of oxygen and CO₂.

[0013] They operate at temperatures of 600°C with a large efficiency of fuel over power production [9]. 5) Alkalines which were used in space applications and produced power with an efficiency of approximately 70% operating at temperatures 60-150°C.

Description of Drawings

[0014]

Figure 1: Schematic diagram of the process of hydrogen and electrical energy production from biomass with a zero pollutant emission. The first part of the process concerns the production of ethanol from biomass and the separation of a fraction of water so as to obtain a 50-60% by weight ethanol concentration. The second part concerns the reactions of partial oxidation/reforming and shift of ethanol for the production of hydrogen and its separation from CO₂ if necessary. The third part concerns the production of electric energy using fuel cells.

Figure 2: Schematic diagram of a reactor suitable for ethanol partial oxidation/reforming reactions. The reactor consists of a bundle of ceramic or metallic tubes. In the internal surface of the tubes the partial oxidation catalyst is deposited in the form of a thin film while in the external surface the reforming catalyst is deposited. The reactor gives a high thermal stability and safe operation over high temperatures.

Description of proposed embodiment

[0015] The process of hydrogen and electrical energy production which is invented is presented diagrammatically in Figure 1. The first part of the process concerns the production of ethanol via fermentation of sugars which are obtained directly from biomass or which are produced from cellulosic raw materials of biomass or from any type of suitable biomass which is a byproduct or waste product of industries involved in processing of agricultural or forestry products. The process of fermentation of sugars and production of ethanol is known and is applied industrially in large scale [1,2]. The ethanol which is produced by fermentation is contained in aqueous solution in concentrations of approximately 8-12%. The aqueous solution is then distilled so as to increase the ethanol content to the level of 50-60% by weight. The energy which is required for the distillation is obtained from the partial oxidation/reforming reactor (see

Figure 1). The process of distillation is also well known and is applied in large scale in chemical and petrochemical industries.

[0016] The aqueous solution which contains ethanol is subsequently fed to the partial oxidation/reforming reactor into which air is simultaneously fed with such a flow rate so that the quantities of ethanol and oxygen to be in stoichiometric proportion according to Reaction 2. Of course, the quantities of oxygen (air) which are fed to the reactor may be smaller or larger than the stoichiometric proportion, depending on the derived characteristics of the reaction, mainly the heat of reaction. For the reactions of partial oxidation/reforming of ethanol a suitable catalyst is required which must be sufficiently active and stable in time of use and especially selective for hydrogen production. Such a catalyst may contain metals of Group VIII of the Periodic Table or metals/metal oxides of the Transition Metals or combination of the above, either by themselves or supported onto different carriers which are used for this purpose. A good catalyst for the partial oxidation/reforming of ethanol is a catalyst which contains Ni dispersed onto a Lanthanum oxide carrier, which is described in a Patent Application [10]. Another good catalyst consists of metals of Group VIII of the Periodic Table dispersed on a carrier of Titanium Oxide doped with cations of Tungsten (W) [11].

[0017] High selectivity towards CO_2 in the reforming reaction is desirable since at least some types of fuel cells do not operate efficiently when the feed contains CO at concentration higher than 20-100 ppm, due to poisoning of the platinum electrode. Because the concentration of CO at the exit of the partial oxidation/reforming reactor will be higher than 100 ppm, processing of the gas in a shift reactor is required in which the following reaction is taking place:



[0018] The shift reaction is an equilibrium reaction and the conversion of CO increases with reduced temperature. This reaction has been studied by many researchers and there are various efficient catalysts for it [2]. If at the exit of the shift reactor the concentration of CO is higher than 100 ppm, this must be oxidized in the presence of a suitable catalyst, in another reactor in which a small quantity of oxygen or air is fed. Gold catalysts supported on selected carriers are active in the selective oxidation of CO in the presence of hydrogen at low temperatures [13]. There is also the alternative solution of hydrogenation of CO towards CH_4 , a reaction which is easily conducted over numerous catalysts.

[0019] The reaction of partial oxidation/reforming of ethanol can take place in any type of suitable reactor, as for example fixed bed reactor, fluidized bed reactor etc. For reactions of this type whose one part is exothermic (partial oxidation) and the other endothermic (reforming) we have developed a particularly efficient and safe reactor which consists of a bundle of ceramic or metallic tubes of small diameter which are encased in a vessel which is thermally insulated, as presented in Figure 2. In the internal area of the tubes the partial oxidation catalyst is deposited while in the external area the reforming catalyst is deposited. The two catalysts may be the same or different. The heat which is produced by the oxidation of part of the fuel is transported through the wall in the external area of the tube, where the endothermic reforming reactions are taking place which consume the transported heat. In this way the reactor can operate adiabatically, the temperature distribution along the length of the reactor is controlled to a large extent, the operation of the reactor is safe since no hot zones are developed, the danger of explosion is eliminated, the pressure drop is small and the productivity of the reactor is high due to the optimization of the heat management of the reaction.

[0020] As is shown in Figure 2, a large fraction of the length of the tubes does not contain catalysts. This part of the tubes is used as heat exchanger to heat the feed, which is introduced at low temperatures, from the reaction products which are at high temperature. In this manner the total volume of the process is minimized.

[0021] The shift reaction can take place in the same or different reactor. In the first case, in the external area of the tubes a catalyst for this reaction is deposited. Alternatively, catalysts in particle or pellet form can be placed in the volume external to the tubes if a larger area is required.

[0022] The length of the tubes which contains catalysts in the internal and external surface, the quantity of catalysts on each side, the type or types of catalysts, the total length of the tubes, the space at which the shift reaction is taking place, the diameter of the encasing vessel, and other similar parameters are defined in such a way so as to maximize the productivity of the reactor.

[0023] The reaction of partial oxidation/reforming of ethanol, with the stoichiometry described by Equation 2, is exothermic and for every mole of ethanol which reacts approximately 20 kcal of heat are produced. Of course, the exact quantity of heat which is produced can be controlled depending on the quantity of air or oxygen which is fed into the reactor. The heat of this reaction can be used not only for the preheating of the reactants (along with the heat of the products of the reaction) but also in the process of distillation of the aqueous solution of ethanol from the initial concentration of 8 to 12% to the desired concentration of 50 to 60%.

[0024] The hydrogen which is produced with the series of reactions which were described above i.e. the partial oxidation/reforming of ethanol, the shift reaction and, if necessary, the reaction of selective oxidation of CO, is fed to

a fuel cell in which it is being oxidized electrochemically with oxygen with simultaneous production of electrical energy and heat. The most suitable fuel cells for the present application are those of the phosphoric acid type or the proton exchange membrane or solid polymers. The heat which is produced simultaneous with the electrical energy is generally of low quality because of the low temperatures (100-200°C) and can be used for the heating of greenhouses if the

5 application is taking place in an agricultural area or buildings if the application is taking place near an urban center.

[0025] The process for production of hydrogen and electrical energy which is described above and is presented diagrammatically in Figure 1 is a typical such process in which many alterations can take place, depending on specific situations or specific applications, without altering the present invention. The same is true for the catalysts and- reactors of the partial oxidation/reforming of ethanol.

10 **Citations**

[0026]

- 15 1. C.E. Wyman (Editor), "Handbook on Bioethanol: Production and Utilization", Taylor & Francis, 1996.
2. D.O. Hall, G. Grassi and H. Scheer (Editors), "Biomass for Energy and Industry", Proc. 7th Int. Conf. Biomass, Florence, Italy, 5-9 October 1992.
3. N. Takezawa and N. Iwasa, Catal. Today, 36 (1997) 45.
4. M. Morita, A. Sato and Y. Matsuda, Nippon Kagaku Kaishi, (1993) 164.
- 20 5. D. Wang, D. Montane and E. Chornet, Appl. Catal. A, 143 (1996) 245.
6. R. Friberg, Int. J. Hydrogen Energy, 18 (1993) 853.
7. V. Recupero, V. Alderrucci, R. Di Leonardo, Int. J. Hydrogen Energy, 19 (1994) 633.
8. T.R. Ralph, Platinum Metals Rev., 41 (1997) 102.
9. Q.M. Nguyen, Chemtech, January 1992, p. 32.
- 25 10. Z.L. Zhang and X.E. Verykios, "A stable and active nickel catalyst for carbon dioxide reforming of methane to synthesis gas", European Patent Application 94600005.6-2104/13.07.94.
11. P. Papaefthimiou, Th. Ioannidis and X. Verykios, "Catalysts for the Combustion of Volatile Organic Compounds (VOC)", Greek Patent Application 970100021/24.1.1997.
12. F. Gottschalk, Appl. Catal., 51 (1994) 127.
- 30 13. M. Haruta, Catal. Surveys of Japan, 1 (1997) 61.

Claims

35 1. A process for producing hydrogen and electrical energy with no emission of pollutants from ethanol obtained from a biomass which comprises the steps of:

(a) providing an aqueous solutions of ethanol in a concentration of approximately 8 to 12% originating from fermentation of a biomass;

40 (b) separating water from the aqueous solution of ethanol so that the aqueous solution of ethanol contains approximately 40 to 70% by weight ethanol;

(c) mixing a portion of the aqueous solution of ethanol obtained according to step (b) with air or oxygen to form a mixture in such a way that the ratio of moles of oxygen per mole of ethanol is between 0.0 and 0.5;

45 (d) feeding the aqueous ethanol obtained according to step (b) and the mixture containing aqueous ethanol and air or oxygen according to step (c) to a thermally insulated reactor comprising a thermally insulated casing and ceramic or metal tubes comprising a tube wall in the internal area of which and in the external area of which catalysts are placed which catalyze reforming ethanol to obtain a gaseous mixture containing carbon dioxide and hydrogen and which catalyze oxidizing ethanol to obtain water, carbon dioxide, hydrogen with liberation of heat wherein each of the catalysts is the same or different;

50 (e) adiabatically reforming and partially oxidizing respectively the aqueous ethanol and the mixture containing aqueous ethanol and air or oxygen in the thermally insulated reactor wherein the heat of reaction needed to carry out the reforming of the aqueous ethanol is supplied by the partial oxidizing of the mixture containing aqueous ethanol and air or oxygen and is transported through the walls of the tubes within the thermally insulated reactor from the partial oxidation reaction to the reforming reaction to obtain the gaseous mixture containing carbon dioxide and hydrogen and to control temperature distribution along the length of the thermally insulated reactor;

55 (f) following the reforming of the aqueous ethanol in step (e) to obtain the gaseous mixture containing carbon dioxide and hydrogen, either performing a water gas shift reaction to convert any carbon monoxide formed in

the gaseous mixture during the reforming along with water vapor to hydrogen and carbon dioxide, or combusting the carbon monoxide, or converting the carbon monoxide to methane; and
 (g) separating the hydrogen from the gaseous mixture in the thermally insulated reactor, and feeding the hydrogen to a fuel cell for producing electrical energy.

5 2. The process for producing hydrogen defined in claim 1 wherein according to step (g) the hydrogen is fed as fuel to a fuel cell for the production of electrical energy and heat.

10 3. The process for producing hydrogen defined in claim 1 wherein according to step (g) the fuel cell is of the phosphoric acid type or proton exchange membrane or solid polymer.

15 4. The process for producing hydrogen defined in claim 1 wherein according to step (d) the catalyst is a Group VIII metal or a transition metal oxide, either supported on a carrier or not.

20 5. The process for producing hydrogen defined in claim 1 wherein according to step (b) the aqueous solution of ethanol contains 50 to 60% by weight ethanol.

Patentansprüche

20 1. Verfahren zur Produktion von Wasserstoff und von elektrischer Energie, ohne Emission von Schmutzstoffen, aus Ethanol, das aus Biomasse gewonnen wurde. Das Verfahren besteht aus den folgenden Stufen:

25 a) Zugabe einer wäßrigen Lösung von Ethanol in einer Konzentration von etwa 8 bis 12%, wobei die Lösung aus der Fermentation einer Biomasse hervorgegangen ist;

30 b) Abtrennung des Wassers aus der wäßrigen Ethanol-Lösung, so daß die wäßrige Ethanol-Lösung etwa 40 bis 70 Gew.-% Ethanol enthält;

35 c) Mischung eines Teils der gemäß Stufe (b) erhaltenen wäßrigen Ethanol-Lösung mit Luft oder Sauerstoff zur Bildung einer Mischung auf solcher Art und Weise, daß das Verhältnis der Sauerstoff-Moleküle pro Ethanol-Molekül zwischen 0,0 und 0,5 beträgt;

40 d) Zufuhr der gemäß Stufe (b) erhaltenen wäßrigen Ethanol-Lösung und der gemäß Stufe (c) Mischung, die wäßriges Ethanol und Luft oder Sauerstoff enthält, zu einem thermisch isolierten Reaktor bestehend aus einem thermisch isolierten Gehäuse und aus keramischen oder metallischen Rohren mit einer Rohrwand, an deren Innenseite und an deren Außenseite Katalyten angeordnet werden, die das Reformieren des Ethanols zu einer Gasmischung, die Wasserstoff enthält, katalysieren und die die Oxidation des Ethanols zu Wasser, Kohlendioxid, Wasserstoff mit Wärmeentwicklung katalysieren, wobei die Katalyten gleich oder verschieden sind;

45 e) adiabatisches Reformieren und Teiloxidation des wäßrigen Ethanols bzw. Mischung aus wäßrigem Ethanol und Luft oder Wasserstoff im thermisch isolierten Reaktor, wobei die Reaktionswärme, die zum Reformieren des wäßrigen Ethanols benötigt wird, aus der Teiloxidation der Mischung aus wäßrigem Ethanol und Luft oder Sauerstoff gewonnen wird und durch die Rohrwände im thermisch isolierten Reaktor von der Teiloxidationsreaktion zur Reformationsreaktion transportiert wird, so daß die Gasmischung, die Kohlendioxid und Wasserstoff enthält, gewonnen wird und so daß die Wärmeverteilung in Längsrichtung des thermisch isolierten Reaktors kontrolliert wird;

50 f) nach dem Reformieren des wäßrigen Ethanols in Stufe (e) zur Erhaltung der Gaslösung, die Kohlendioxid und Wasserstoff enthält, Durchführung eines Wassergasverfahrens zur Umsetzung von Kohlenmonoxid, das sich in der Gasmischung beim Reformieren zusammen mit Wasserdampf gebildet hat, zu Wasserstoff und Kohlendioxid, oder Verbrennung des Kohlenmonoxids oder Umsetzung des Kohlenmonoxids zu Methan, und

55 g) Abtrennung des Wasserstoffs aus der Gasmischung im thermisch isolierten Reaktor, und Zufuhr des Wasserstoffs zu einer Brennstoffzelle zur Produktion von elektrischer Energie.

2. Verfahren zur Produktion von Wasserstoff nach Anspruch 1, wobei gemäß Stufe (g) der Wasserstoff als Brennstoff einer Brennstoffzelle zur Produktion von elektrischer Energie und Wärme zugeführt wird.

3. Verfahren zur Produktion von Wasserstoff nach Anspruch 1, wobei gemäß Stufe (g) die Brennstoffzelle vom Phosphorsäure-Typ oder eine Protonenaustauschmembran oder ein festes Polymer ist.

4. Verfahren zur Produktion von Wasserstoff nach Anspruch 1, wobei gemäß Stufe (d) der Katalyt ein Metall der Gruppe VIII oder ein Übergangsmetallocid ist, entweder unterstützt von einem Träger oder nicht.

5. Verfahren zur Produktion von Wasserstoff nach Anspruch 1, wobei gemäß Stufe (b) die wäßrige Ethanol-Lösung 50 bis 60 Gew. -8% Ethanol enthält.

5 **Revendications**

1. Procédé de production d'hydrogène et d'énergie électrique, sans émission polluante, à partir d'éthanol produit par la biomasse, qui comprend les étapes suivantes:

10 a) provision d'une solution aqueuse d'éthanol ayant une concentration entre environ 8 et 12%, produite par la fermentation de la biomasse;

b) séparation d'eau de la solution aqueuse d'éthanol pour que la solution aqueuse d'éthanol contienne d'éthanol entre environ 40 et 70% en poids;

c) mélange d'une portion de la solution aqueuse d'éthanol obtenue selon l'étape (b) avec l'air ou l'oxygène pour former un mélange où le rapport molaire oxygène/éthanol est entre 0,0 et 0,5;

15 d) alimentation de l'éthanol aqueux obtenu selon l'étape (b) et du mélange qui contient de l'éthanol aqueux et de l'air ou l'oxygène selon l'étape (c) dans un réacteur thermiquement isolé comprenant une enveloppe thermiquement isolée et des tubes céramiques ou métalliques, comprenant une paroi des tubes, sur la surface extérieure de laquelle et la surface intérieure de laquelle catalyseurs sont placés, qui catalysent le reformage de l'éthanol pour obtenir un mélange gazeux contenant du dioxyde de carbone et de l'hydrogène et qui catalysent l'oxydation de l'éthanol pour obtenir l'eau, le dioxyde de carbone, l'hydrogène avec la libération de chaleur, où chacun catalyseur est le même ou différent;

20 e) reformage adiabatique et oxydation partielle respectivement de l'éthanol aqueux et du mélange contenant de l'éthanol aqueux et de l'air ou l'oxygène dans le réacteur thermiquement isolé où la chaleur de la réaction qui est nécessaire pour réaliser le reformage de l'éthanol aqueux est fournie par l'oxydation partielle du mélange contenant de l'éthanol aqueux et de l'air ou l'oxygène et est transportée à travers les parois des tubes dans le réacteur thermiquement isolé, de la réaction d'oxydation partielle à la réaction du reformage pour obtenir le mélange gazeux contenant du dioxyde de carbone et de l'hydrogène et pour contrôler la distribution de température tout le long du réacteur thermiquement isolé;

25 f) après le reformage de l'éthanol aqueux à l'étape (e) pour obtenir le mélange gazeux contenant du dioxyde de carbone et de l'hydrogène réalisation d'une réaction de déplacement de gaz à l'eau pour convertir le monoxyde de carbone formé dans le mélange gazeux pendant le reformage, par la vapeur d'eau en hydrogène et dioxyde de carbone, ou la combustion du monoxyde de carbone ou la conversion du monoxyde de carbone en méthane; et

30 g) séparation de l'hydrogène du mélange gazeux dans le réacteur thermiquement isolé, et alimentation de l'hydrogène dans une pile à combustible pour la production d'énergie électrique.

2. Procédé de production d'hydrogène selon la revendication 1 où selon l'étape (g) l'hydrogène est alimenté comme combustible dans une pile à combustible pour la production d'énergie électrique et de chaleur.

3. Procédé de production d'hydrogène selon la revendication 1 où selon l'étape (g) la pile à combustible est une pile à acide phosphorique ou à membrane échangeuse de protons ou à polymères solides.

4. Procédé de production d'hydrogène selon la revendication 1 où selon l'étape (d) le catalyseur est un métal du groupe VIII ou un oxyde d'un métal de transition, supporté ou non-supporté sur un matériau de support.

45 5. Procédé de production d'hydrogène selon la revendication 1 où selon l'étape (b) la solution aqueuse d'éthanol contient d'éthanol entre 50 et 60% en poids.

50

55

Collection of Biomass

Production of Bio-ethanol

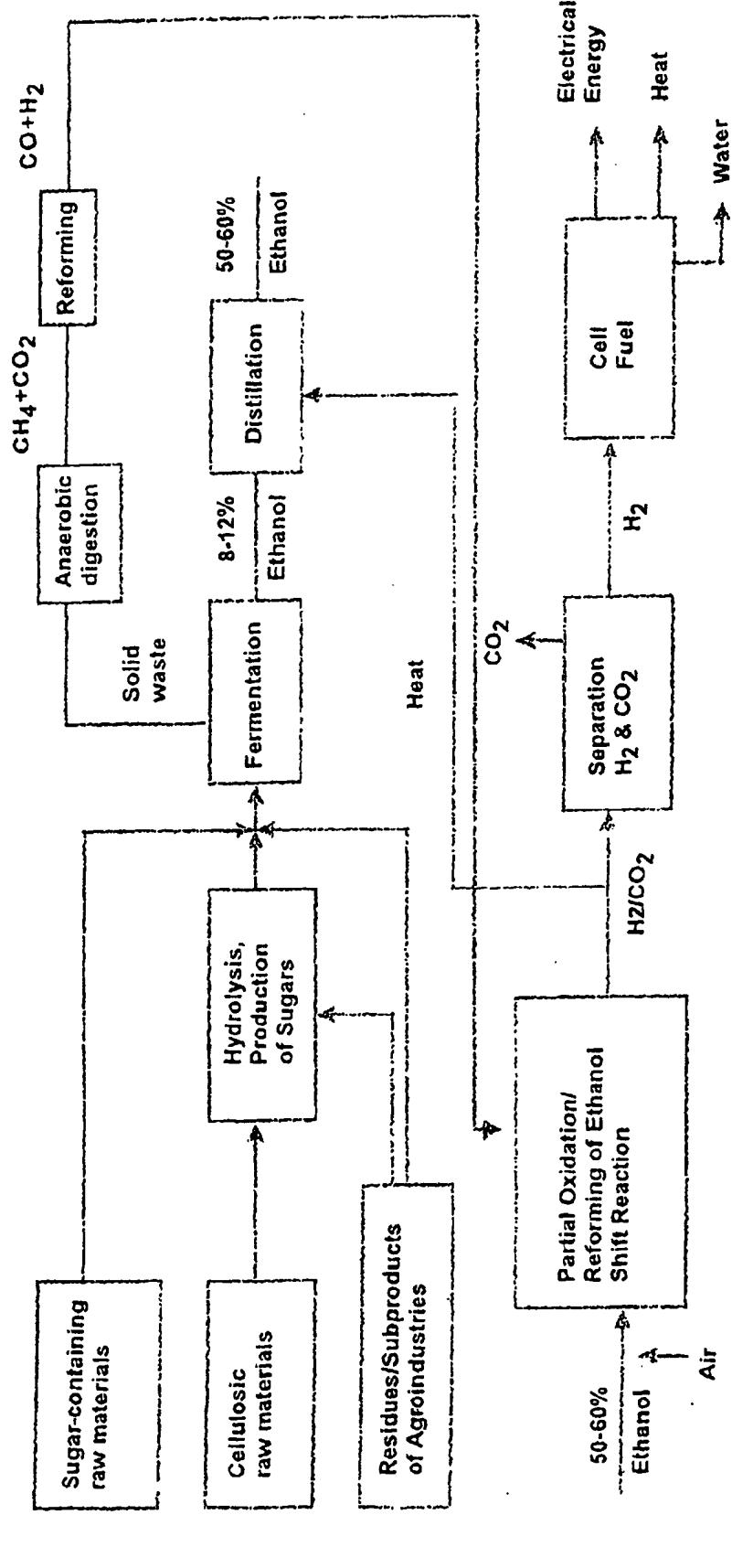


Figure 1

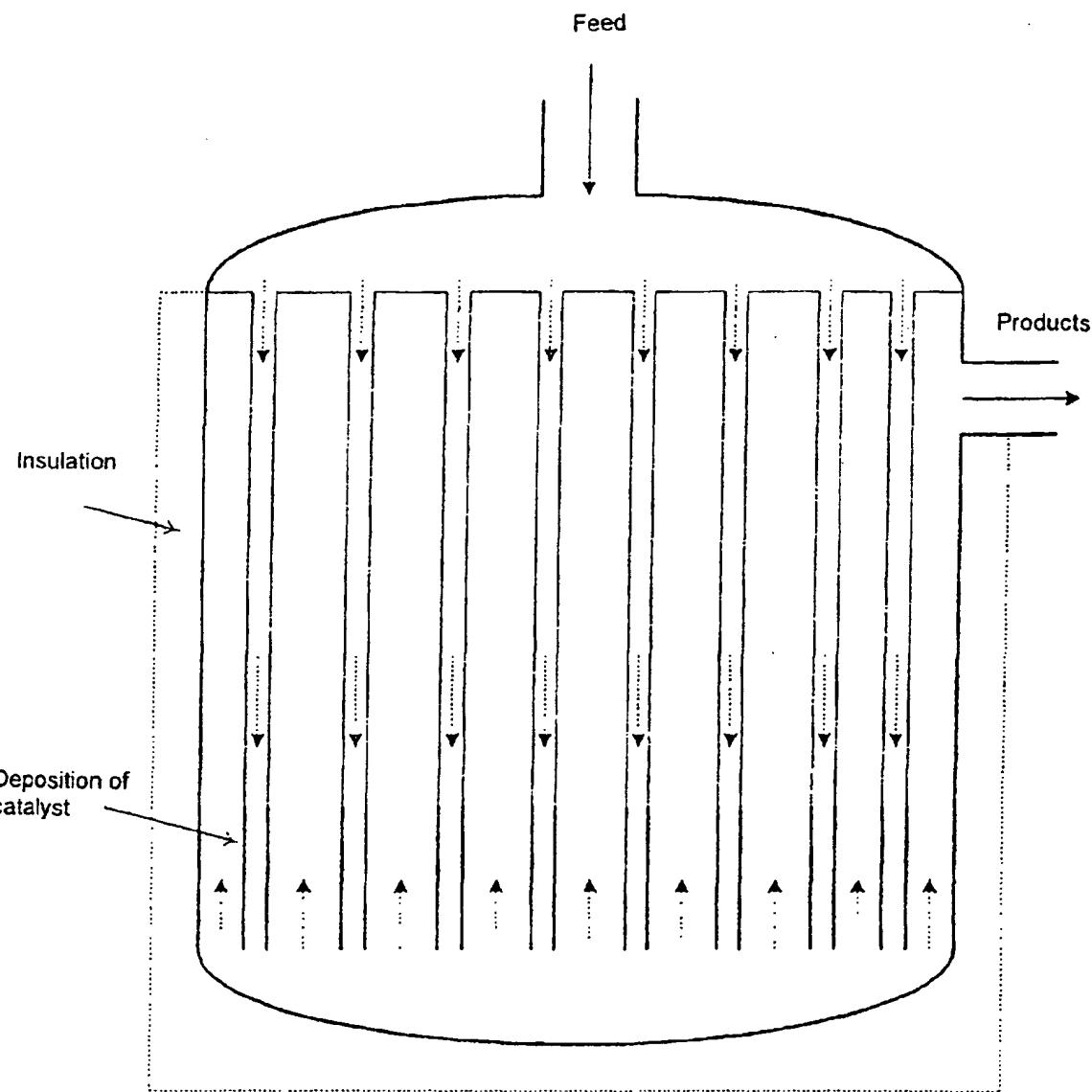


Figure 2